

MOISEYEV, L.L., aspirant

Regulation of the productive capacity of piston compressors.

Sbor. nauch. trud. Kem. gor. inst. no. 5:144-149 '64.

(MIRA 18:3)

1. Gorno-elektromekhanicheskiy fakul'tet Kemerovskogo gornogo instituta.

APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6

PANKRAT'YEV, Vladimir Pavlovich; TOMILIN, Yu.K.; MOISEYEV, L.K.:
KOSTINSKIY, D., red.

[United Republic of Tanzania] Ob"edinennaia Respublika
Tanzania. Moskva, Mysl', 1965. 94 p. (MIRA 18:4)

MOISEYEV, L.K.

Zanzibar. Geog. v shkole 26 no.3:67-74 My-Je

(MIRA 16:6)

(Zanzibar--Description and travel)

SALYUKOV, P.A., kand. biol. nauk; VERNIGOR, V.A., kand. sel'khoz. nauk; KORMANOVSKAYA, M.A., kand. sel'khoz. nauk; GOLODNOV, A.V.; SKOROBOGATOV, Yu.A., mladshiy nauchnyy sotr.; MALLITSKIY, V.A., kand. sel'khoz. nauk; CHASHCHIN, B.V., kand. sel'khoz. nauk; PONOMAREV, P.P., kand. tekhn. nauk; BARMINTSEV, Yu.N., doktor sel'khoz. nauk; NECHAYEV, I.N., mlad. nauchnyy sotr.; POZDNYAKOV, P.M., kand. biol. nauk; KOVIN'KO, D.A., kand. biol. nauk; BALANINA, O.V., kand. sel'khoz. nauk; MOISEYEV, K.V., kand. sel'khoz. nauk; ROMANOV, P.F., kand. veter. nauk; PAL'GOV, A.A., kand. veter. nauk; ANAN'YEV, P.K., kand. veter. nauk; VASIL'YEV, B.M., kand. sel'khoz. nauk; ABDULLIN, V.A., kand. ekon. nauk; GALIAKBEROV, N., laureat Gos. premii, kand. sel'khoz. nauk, red.; GUSEVA, N., med.; NAGIBIN, P., tekhn. red.

[Reference book for zootechnicians] Spravochnik zootekhnika.
Pod red. N.Galiakberova. Alma-Ata, Kazsel'khozgiz, 1963.
492 p. (MIRA 16:5)
(Kazakhstan--Stock and stockbreeding)

USSR/Fern Animals. Honey Bee

C-6

Abstr Jour : Ref Zhur - Biol., No 8, 1958, No 35767

Author : ~~Moiseyev K.V.~~

Inst : Not Given

Title : Crossbreeding of the Local Honey Bees of Kazakhstan with the Mountain Caucasian Breed (Skroshchivaniye mostnykh pchel Kazakhstana s gornymi kavkazskimi)

Orig Pub : Pchelopodstvo, 1957, No 9, 24-29; See RZhBiol, 1957, No 24, ref. 103357

Abstract : No abstract

Cord : 1/1

1. MIN'KOV, D. G.; NOTSEYEV, R. V.
2. USSR (600)
4. Parasites - Bees
7. Control of blister beetle larvae [eloe]. Zheleznodolno, 30, No. 4, 1953.

9. Monthly List of Russian Accessions, Library of Congress, _____ April _____ 1953, Uncl.

MOISEYEV, K. V.

Bee Culture - Queen Rearing

Time for shipping queen bees Pchelovodstvo 29, no. 4, April 1952

9. Monthly List of Russian Accessions, Library of Congress, August ² 1973, Unc .

MOISEYEV, K.N.

Fisheries - Accounting

Putting the shop on a completely self-financing basis. Ryb. Khoz. 28 no. 11, 1952.

Monthly List of Russian Accessions. Library of Congress. November 1952. Unclassified

MOISEYEV, K.I., student; SIVUKHA, M.I., student; SHUL'KIN, Yu.P., student

Investigating the endurance limit of shipbuilding steel to bending
on a level plane together with axial deformation. Trudy LKI no.29:
221-224 '59. (MIRA 14:7)

1. Leningradskiy korablestroitel'nyy institut, korablestroitel'nyy
fakul'tet.

(Plates, Iron and steel--Testing)

MOISEYEV, K. I.

MOISEYEV, K. I. "On the methodology of reconstructing the conjunctiva", Sbornik nauch. trudov vrachey Mordov. ASSR, Saransk, 1948, p. 47-49

SC: U-3.61, 10 April 53 (Letopis - Zhurnal 'nykh Stat'ey No. 11, 1949)

APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6

VAVILOV, I. P., 4011 1st Avenue, N.E., Seattle, Wash., U.S.A.
 H. ISENH, 6, rue de la Chapelle, Paris, France

[illegible]

VAVILOV, P.P.; MOISEYEV, K.A.

Introduction of silage plants and their propagation on state and
collective farms of the Komi A.S.S.R. Biol.Glav.bot.sada no. 48:
3-11 '63. (MIRA 17:5)

1. Komi filial AN SSSR, gorod Syktyvkar.

VAVILOV, P.P., kand. sel'khoz. nauk, glav. red.; LAZAREV, N.A.,
kand. sel'khoz. nauk, zam. glav. red.; GALAS'YEV, V.A.,
red.; MOISEYEV, K.A., kand. biol. nauk, red.;
PODOPLELOV, V.P., kand. ekon. nauk, red.; STARKOVA, V.N.,
kand. biol. nauk, red.; TARASENKOV, G.H., kand. geogr.
nauk, red.; TON, D.S., kand. ekon. nauk, red.; TIKHONOVA,
N.V., red.isd-va; VDOVINA, V.M., tekhn. red.

[Forests and the lumbering industry in the Komi A.S.S.R.]
Lesa i lesnaia promyshlennost' Komi ASSR. Moskva, Gos-
lesbunizdat, 1961. 394 p. (MIRA 16:4)

1. Akademiya nauk SSSR. Komi filial, Syktyvkar.
(Komi A.S.S.R.--Forests and forestry)

MOISEYEV, K.A.; ALEKSANDROVA, M.I.

Effect of presowing treatment of seeds with trace elements on silage
plants and vegetables. Trudy Komi fil. AN SSSR no.9:23-32 '66.

(TRACE ELEMENTS) (FORAGE PLANTS) (VEGETABLES) (MIRA 19:1)

Country :
 Category : CULTIVATED PLANTS, FRUITS
 Abs. Jour. : DFF ZHUS PIGL. 21.1958, NO 96145

Author :
 Institut. :
 Title :

Orig. Pub. :

Abstract : penetrate the soil deeper than 50 cm. The basic most active part of the root system (up to 75%) occurred in relation to soil moisture 20-40 cm deep. The adventitious root system surpassed the basic roots in weight by 40-80%. In the fifth year of life the adventitious root number began to decrease, the harvest was also considerably reduced and the berries turned tiny. In the fourth year the branches bearing berries produced the greatest yields. It is recommended that the black currant

Card: 2/3

MOISEYEV, K.A.

Country : USSR

Category : CULTIVATED PLANTS, FRUITS, Berries.

Abstr. Jour. : ILEFZHOE-RUSL (2) 1968, NO 9, 165

Author : ~~MOISEYEV, K.A.~~

Co-auth. : Kozlov, A.A. and G. V. P.

Title : Growth and Development of Black Currants Cultivated in the North

Orig. Pub. : Tr. Pom. SSSR, 1968, No. 6, 127-133

Summary : A study was made of the growth and development of the root system of the basic root system of the main and bush and adventitious roots of the basic branches of the black currant, the growth of the adventitious roots, beginning of fruiting, the dynamics of productivity on branches of different rows and the basic branch in relation to age. The Priozorski, Zhuravka, Miya Plodnorodnaya, Slava Petrovskaya and Yuzovka Sve-de varieties were under observation. The root system did not

Card: 1/1

AUTHOR: Moiseyev, K. 307-1-10-9-1/85

TITLE: Attention ! There Are Oil-Fields Underneath ! (Vnimaniye !
V glubine zalegayet neft !)

PERIODICAL: Znaniye-sila, 1958, Nr 8, p 1 (USSR)

ABSTRACT: Soviet oil-prospecting geologists are now using nuclear physics and electronics. A team of young scientists from the Professor Fedor Aleksyevich **Alekseyev** Laboratory at the Institut nefti Akademii nauk SSSR (Oil Institute of the USSR Academy of Sciences) has developed a gamma ray radiometer, used to locate oil deposits. Candidate of Geological and Mineralogical Sciences, V.I. Yermakov asserts that the new gamma ray radiometer will be more sensitive than former devices and can be used in airplanes and submarines. There is 1 drawing.

1. Geophysical prospecting--USSR 2. Radiometers--Applications
3. Nuclear physics--Applications 4. Electronic equipment--Applications
5. Petroleum--Production

Card 1/1

AUTHOR: Moiseyev, K. SOV-4-58-7-12/22

TITLE: The First Peaceful Function of Tritium (Pervaya mirnaya professiya tritiya)

PERIODICAL: Znaniye - sila, 1958, Nr 7, p 23 (USSR)

ABSTRACT: Tritium - an isotope of hydrogen - has till now been used only for military purposes. Recently Soviet scientists discovered a possible peaceful function for tritium. The movement of sub-surface water had not been investigated satisfactorily until a number of scientists of the laboratory of the Institute of Petroleum of the USSR Academy of Sciences headed by Member-Correspondent G.N. Flerov and Professor F.A. Alekseyev successfully used tritium as an "indicator" of the movement of the water below the petroleum layer. In 1956, tritium was applied for the first time for "marking" water by the Grozny oil industry. There is one diagram.

Card 1/1

MOISEYEV, I.Ye.; POPLAVSKIY, S.D.

Widening of the working band of frequencies of the generator and indicator of the KIPZ-300 instrument kit. Vest. svyazi 23 no.7: 17-18 J1 '63. (MIRA 17:2)

1. Nachal'nik laboratorii TSentral'nogo nauchno-issledovatel'skogo instituta svyazi Ministerstva svyazi SSSR (for Moiseyev). 2. Starshiy inzh. TSentral'nogo nauchno-issledovatel'skogo instituta svyazi Ministerstva svyazi SSSR (for Poplavskiy).

MOISEYEV, I.Ye.; POPLAVSKIY, S.D., starshiy inzh.

Widening of the operational frequency band of the ZG-12 generator.
Vest. svyazi 21 no.11:20-21 N '61. (MIRA 14:11)

1. Nachal'nik laboratorii TSentral'nogo nauchno-issledovatel'skogo
instituta svyazi Ministerstva svyazi SSSR (for Moiseyev).
(Oscillators, Electric)

MOISEYEV, I.Ye., inzh., mladshiy nauchnyy sotrudnik.

Controlling device for level indicators and wide-band electronic
voltmeters. Vest.svyazi 18 no.1:10-11 Ja '58. (MIRA 11:1)

1. TSentral'nyy nauchno-issledovatel'skiy institut svyazi.
(Telecommunication--Equipment and supplies)
(Electron-tube voltmeter)

L 41070-65
 ADDITIONAL AN3004022

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 Ch. IV. Weight methods of determining uranium -- 163
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MARKOV, V.K., doktor khim. nauk, prof.; VERNYY, Ye.A., kand. mat. nauk; VINOGRADOV, A.V., kand. khim. nauk; YEFIMOV, S.V., kand. khim. nauk; KLYGIN, A.Ye., kand. khim. nauk; MOISEYEV, I.V., kand. khim. nauk; PANASENKOVA, Ye.L., red.; ALYAB'YEV, A.F., red.

[Uranium; methods for its determination] Urany i metody ego opredeleniia. Izd.2., ispr. i dop. Moskva, Akademiya (MIRA) 1964. 502 p.

19901

S/078/61/006/003/006/022
B121/B208

Investigation of some...

extraction of plutonium(IV) cupferronate from sulfuric acid solutions with chloroform was studied. The composition of the extractable complex is $\text{Pu}(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_4$. Extraction of trivalent plutonium by chloroform from sulfuric and nitric acid solutions in the presence of cupferron also takes place in the form of $\text{Pu}(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_4$. Plutonium(IV) cupferronate is less extracted with CCl_4 than with CHCl_3 . At $20 \pm 1^\circ\text{C}$, the following equation holds for the constant K_{eq} in the system $\text{PuR}_4 - \text{HR} - \text{CHCl}_3$:

$$K_{\text{eq}} = \frac{[\text{PuR}_4]_{\text{CHCl}_3}}{[\text{Pu}^{\text{IV}}]_{\text{H}_2\text{O}}} \cdot \frac{[\text{H}^+]^4}{[\text{HR}]_{\text{CHCl}_3}^4} = 1.1 \cdot 10^7.$$

There are 2 figures, 2 tables, and 8 references: 7 Soviet-bloc and 1 non-Soviet-bloc.

SUBMITTED: July 20, 1960

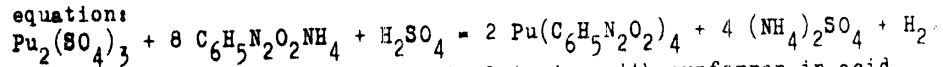
Card 3/3

89901

S/078/61/006/003/006/022
B121/B208

Investigation of some...

valent plutonium with cupferron takes place according to the following equation:



In the precipitation of hexavalent plutonium with cupferron in acid solutions, plutonium(VI) is reduced to plutonium(IV) which is precipitated as $\text{Pu}(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_4$. The solubility of plutonium cupferronate was determined by the method of I. V. Pyatnitskiy (Ref. 6):

$$K_p = \frac{[\text{H}^+]^4}{[\text{Pu}^{\text{IV}}][\text{HR}]^4} = 6.6 \cdot 10^{13}$$

$$L_p = [\text{Pu}^{\text{IV}}][\text{R}^-]^4 = 1.2 \cdot 10^{-31}$$

It was found in numerous experiments that the precipitation of plutonium cupferronate from 1.5 - 3 M sulfuric acid solutions in the presence of 2.5 - 3.5 mg/ml of excess cupferron gives rise to a complete plutonium separation from equimolecular quantities of uranium, chromium, manganese, aluminum, silver, nickel, and lanthanum. Also a quantitative separation of plutonium from americium is achieved under equal conditions. The

Card 2/3

89901

S/078/61/006/003/006/022
B121/B208

213100

AUTHORS: Moiseyev, I. V., Borodina, N. N., Tavetkova, V. T.

TITLE: Investigation of some physico-chemical properties of
plutonium cupferronate

PERIODICAL: Zhurnal neorganicheskoy khimii, v. 6, no. 3, 1961, 543-548

TEXT: The authors studied the composition, solubility, and extraction of the reaction products of tri-, tetra-, and hexavalent plutonium with cupferron, since no data are available on their physico-chemical properties. The composition of the precipitate in the precipitation of plutonium with cupferron was always found to correspond to tetravalent plutonium, irrespective of the plutonium valence in the initial solution. The composition of plutonium(IV) cupferronate was determined by potentiometric titration of sulfuric acid solutions of tetravalent plutonium with cupferron solutions. When precipitating trivalent plutonium with cupferron, the latter is not consumed by oxidation, but the plutonium(III) ion is oxidized in the solution and then forms the stable complex compound $\text{Pu}(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_4$ in the presence of cupferron. Precipitation of tri-

Card 1/3

MOISEYEV, I. V.

Process of the formation of the Lovozero alkaline massif.
Trudy IMGRE no.9:240-247 '62. (MIRA 16:1)

(Lovozero Tundras—Geology, Structural)

Uranium, Methods of Detection

SOV/5117

bulk of the material is on the determination of uranium by gravimetric, volumetric, photometric, electrometric, and radiometric methods. One chapter is devoted to the determination of uranium by the luminescence method. No personalities are mentioned. References accompany each of the chapters.

TABLE OF CONTENTS:

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Ch. I. Physical and Chemical Properties of Uranium and of Some of Its Compounds (V. K. Markov)	5
1. Occurrence of uranium in nature	5
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~~Card 2/11~~

PHASE I BOOK EXPLOITATION

SOV/5117

Markov, V. K., A. V. Vinogradov, S. V. Yelinson, A. Ye. Klygin,
and I. V. Moiseyev

Uran, metody yego opredeleniya (Uranium, Methods of Detection)
Moscow, Atomizdat, 1960. 262 p. Errata slip inserted.
6,000 copies printed.

Ed. (Title page): V. K. Markov, Doctor of Chemical Sciences;
Ed.: Ye. I. Panasenkova; Tech. Ed.: Ye. I. Mazel'.

PURPOSE: This book is intended for technical personnel of the
uranium industry.

COVERAGE: The book contains systematized material on the de-
termination and separation of uranium. Chemical, luminescence,
and radiometric methods for qualitative detection of uranium
in various media are described in detail. The description of
methods for the separation of uranium includes, among others,
precipitation, extraction, and cation and anion exchange. The

Card 1/11

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MCIS : V, I. V.

Kilby; Lev, P. A. and Polak, L. V. "The small
Hole through the wall," *Soviet Science*,
in English, Vol. 1, No. 1, 1968, p. 17-18.

AC: U-3264, of Serial 1003, (Letter to Bureau of Investigation, re: [redacted])

MAKSIMOV, Semen Andreyevich; GOL'TSBERG, I.A., redaktor; MOISEYEV, I.T.,
redaktor; SOLOVYCHIK, A.A., tekhnicheskii redaktor

[Meteorology in agriculture] Meteorologiya i sel'skoe khoziaistvo.
Izd. 3-e, ispr. i dop. Leningrad, Gidrometeorologicheskoe izd-vo,
1955. 141 p. (MLRA 9:1)

(Meteorology, Agricultural)

NOISEYEV, I.S., aspirant

Calculating the thermal regime of earth dams in regions where records
frozen over a period of many years. Spok. trud. MISI no. 24:281-293 '52.
(MIRA 12:7)

(Hydraulic engineering--Cold weather conditions)
(Dams)

10(4), 14(0)

SOV/98-59-9-14/22

AUTHOR: Moiseyev, I.S., and Klimenko, N.G., Engineers

TITLE: Ground Compaction by Self-Propelled Dump Trucks
"MAZ-525"

PERIODICAL: Gidrotekhnicheskoye stroitel'stvo, 1959, Nr 9
p 46 (USSR)

ABSTRACT: The author describes the first Soviet construction of an earth dam compacted exclusively by rolling of self-propelled dump trucks, which serve at the same time for earth transportation. The experiment was carried out during construction of the 12,000,000 cu m earth dam for the Irkutsk GES; for rolling the "MAZ-525" self-propelled dump trucks, the weight of which, when loaded, amounts to 50 tons. The earth was placed in 1-1.5-m layers; a planned system of earth transportation, which enabled an optimum compaction, was applied. There are 1 photograph and 2 graphs.

Card 1/1

MOISEYEV, I.S.

Degradation of the frozen state of earth dams and permafrost
foundations. Nauch.dokl.vys.shkoly; stroi. no.2:229-230 '59.
(MIRA 13:4)

1. Rekomendovana kafedroy gidrotekhnicheskikh sooruzheniy
Moskovskogo inzhenerno-stroitel'nogo instituta imeni V.V.
Kuybysheva.

(Frozen ground) (Foundations) (Dams)

SEVNY, Vladimir Ivanovich; MOISEYEV, I.S., red.; MATVEYEV, G.I., tekhn.red.

[Protecting foundation holes of hydraulic structures against
underground waters] Zashchita kotlovanov gidrotekhnicheskikh
sooruzhenii ot gruntovykh vod. Moskva, Gos.energ.isd-vo, 1959.
88 p. (MIRA 12:5)
(Water, Underground) (Drainage) (Hydraulic engineering)

MOISEYEV, I. S. Cand Tech Sci -- (diss) "Methods of thermal calculation of frozen-earth ^{layers} ~~masses~~." Mos, 1959. 14 pp (Min of Higher and Secondary Specialized Education RSFSR. Mos Order of Labor Red Banner Construction Engineering Inst im V. V. Kuybyshev), 200 copies (KL, 50-59, 127)

MOISEYEV, I.S., aspirant

Heat calculations for lightweight concrete dams to be built in
permafrost areas. Nauch.dokl.vys.shkoly; stroi. no.2:281-286
' 58. (MIRA 12:1)

(Frozen ground) (Dams)

99-1-9/10

Time of Filling of Layers at the Construction of Dams During the Winter Months.

There is 1 photograph and 2 graphs.

AVAILABLE: Library of Congress

Card 2/2

MOISEYEV, I.S.

99-1-9/10

AUTHOR: Moiseyev, I.S., Engineer

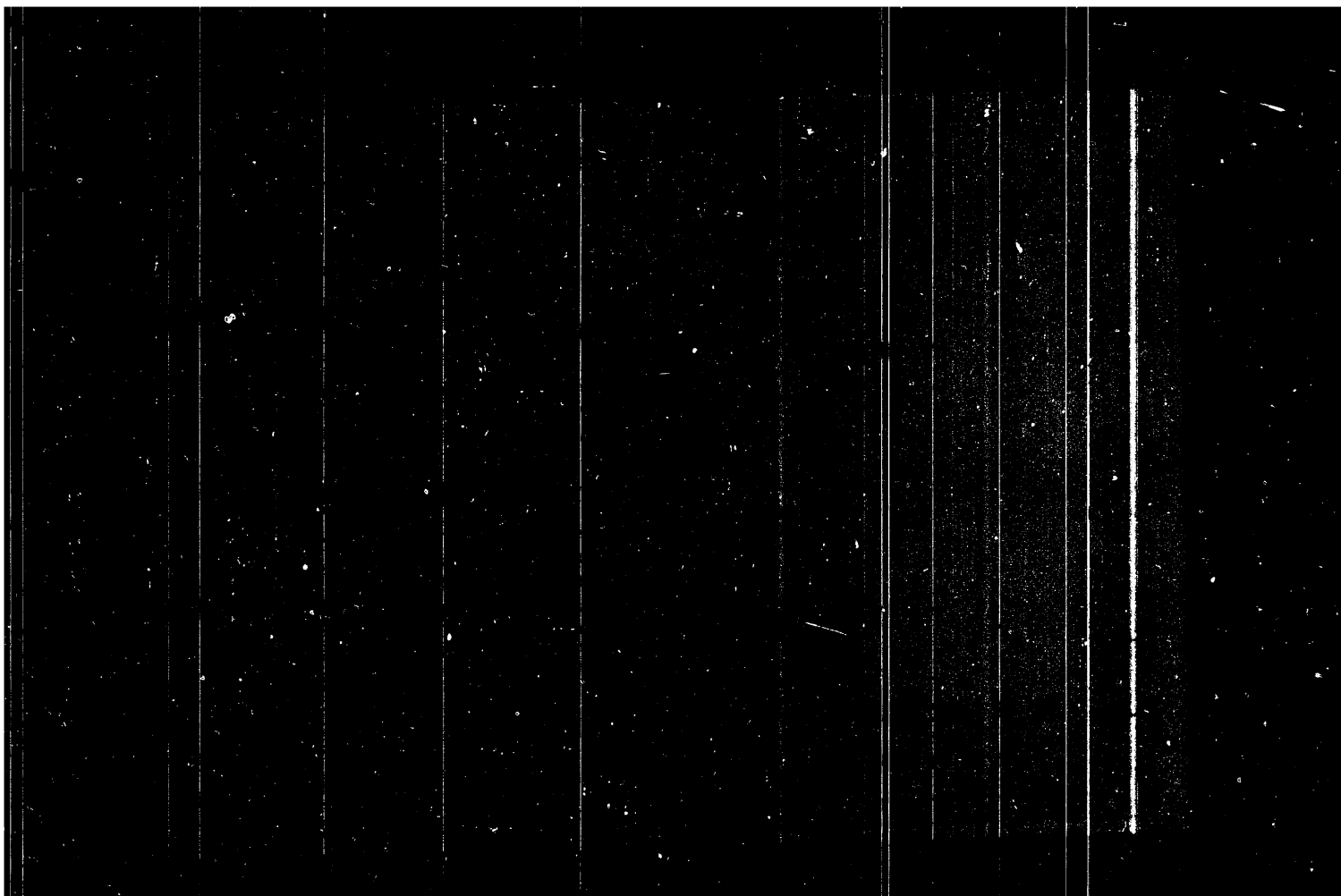
TITLE: Time of Filling of Layers at the Construction of Dams During the Winter Months (O srokakh perekrytiya kart otaypki pri vozvedenii nasypnykh plotin zimoy)

PERIODICAL: Gidrotekhnika i Melioratsiya, 1958, # 1, pp 59-60 (USSR)

ABSTRACT: Earth dams composed of gravel-sand dumped in 1.5 m thick horizontal layers can be built under any climatic conditions. Construction time is shortened and mechanical equipment is used more economically when filling operations are conducted throughout the year. As the filling of one layer generally takes 20 - 30 days, the entire layer is almost completely frozen by the time its respective section is finished. The author examined the process and the effects of freezing for the purpose of establishing the maximum periods for filling individual sections. The author developed several equations and graphs based on various factors, among which are soil humidity, mean air temperatures during the freezing period, etc. It can be seen from the graphs that the depth of freezing at constant air temperature decreases with increased humidity of the ground.

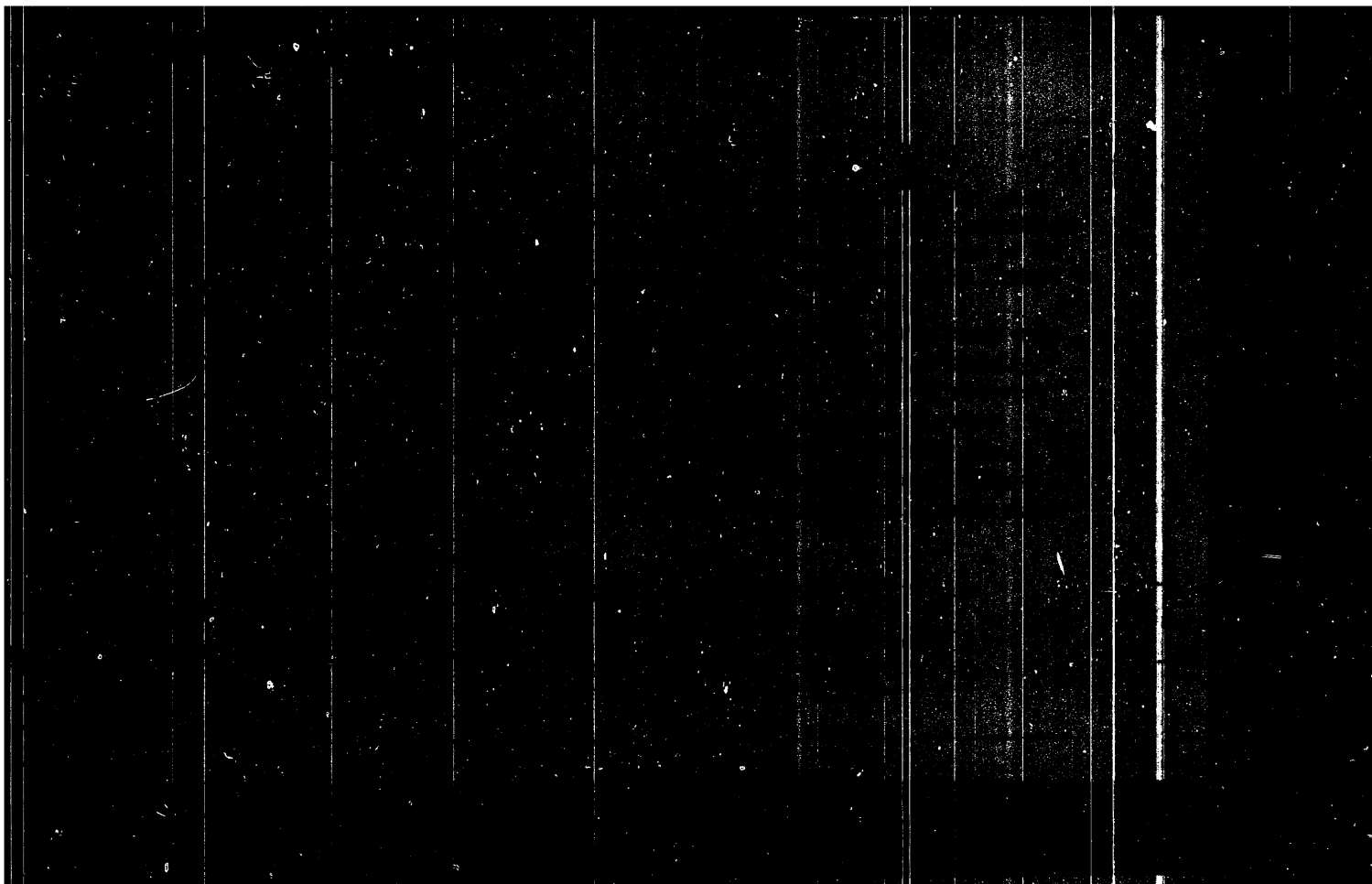
Card 1/2

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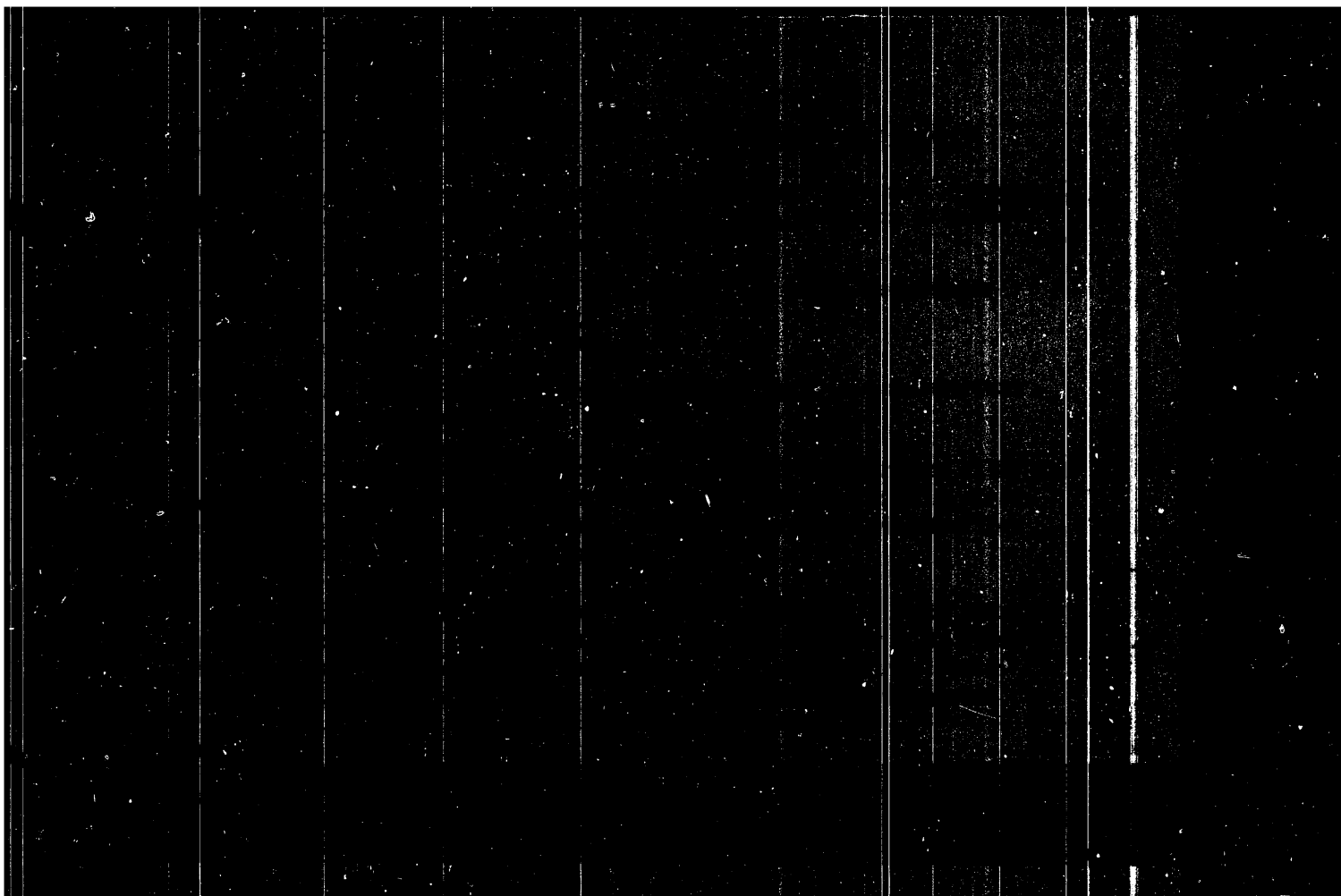
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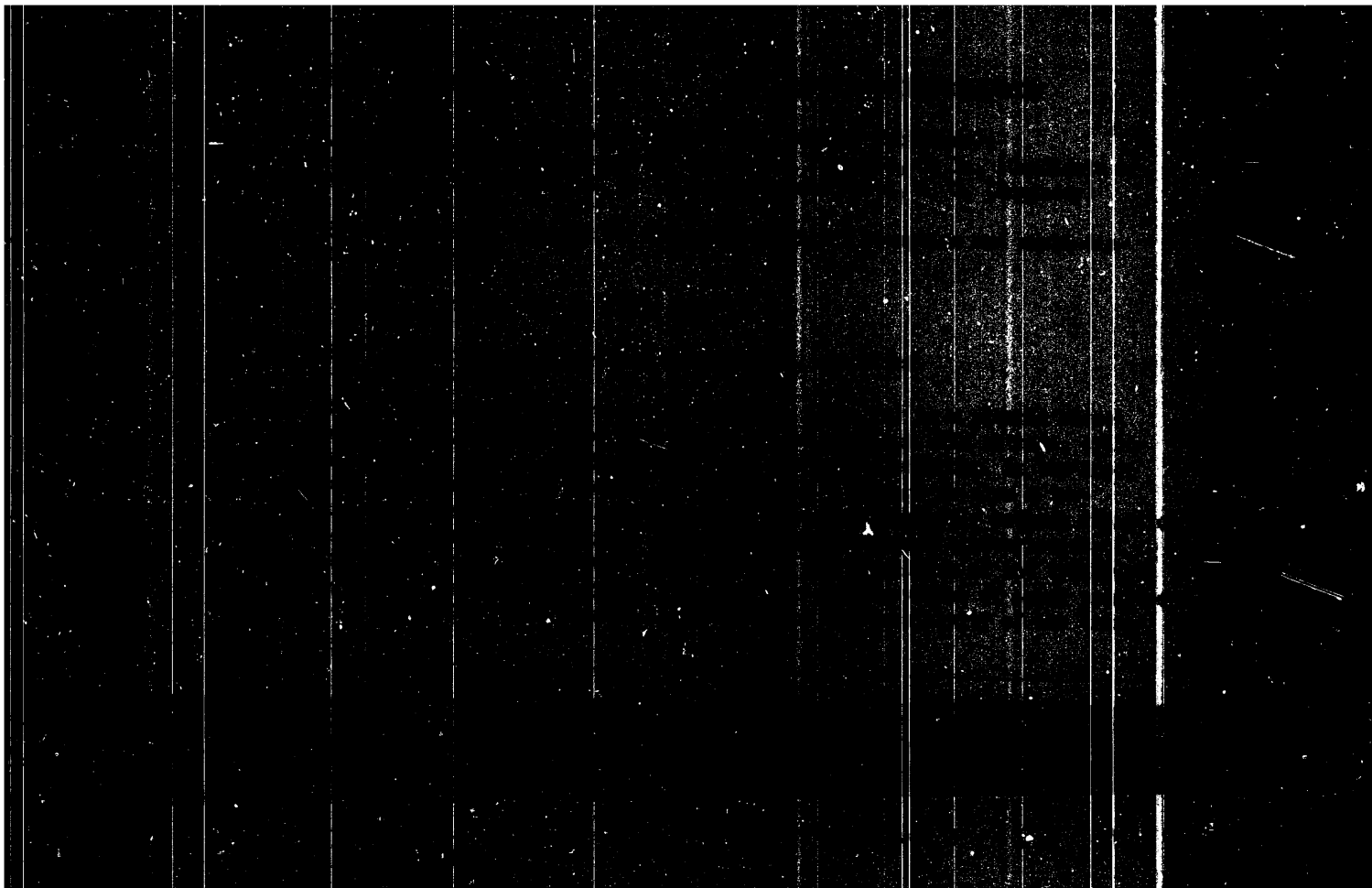


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[Hydrological yearbook] Gidrologicheskii eshegodnik. Lenin-grad, Gidrometeoizdat. 1958. Vol.3. [Basins of the rivers of the Caucasus] Basseiny rek Kavkaza. Nos.2-5. [Basins of the rivers of the Black Sea from the basin of the Kuban River to the frontier and the basin of the Caspian Sea Southward from the basin of the Terek River to the frontier] Basseiny rek Chernogo moria ot basseina r. Kuban' do gosudarstvennoi granitsy i bassein Kaspiiskogo moria k iugu ot basseina r.Terek do gosudarstvennoi granitsy. Pod red. T.N.Dzhaparidze. 1962. 418 p. (MIRA 16:5)

(Hydrology--Tables, calculations, etc.)

KHRAPOVITSKAYA, M.K., red.; MOISEYEV, I.N., red.; VOLKOV, N.V.,
tekhn. red.

[Hydrological yearbook] Gidrologiches kii ezhegodnik. Len-
ningrad, Gidrometeoizdat. 1957. Vol.6.[Basin of the Kara
Sea (western part)] Bassein Karakogo moria (zapadnaya
chast') Nos.0-3. [Ob' River and its basin to the estuary
of the Irtysh River] Reka Ob' i ee bassein do ust'ia r.
Irtysa. Pod red. M.K.Khrapovitskoi. 1962. 325 p.
(MIRA 16:12)
(Ob' Valley--Hydrology--Tables, calculations, etc.)

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2. The following information is being furnished to you for your information:
3. The following information is being furnished to you for your information:

BELOV, A.P.; PEK, G.Yu., MOISEYEV, I.I.

Oxidation of propylene by palladium chloride in acetic acid.
Izv. AN SSSR Ser. khim. no. 12:2704-2706, 1965.

(NINA 18:10)

I. Moskovskiy Institut tekhnicheskoy khimii (Moscow).
Lomonosova 2 Institut khimicheskoy tekhnologii.
N.S. Kurnakova AN SSSR. Submitted April 7, 1965.

PESTRIKOV, S.V.; MOISEYEV, I.J.; ROMANOVA, T.N.

Effect of temperature on equilibrium of the formation of π -complex of palladium chloride with ethylene, propylene, and 1-butylene in water. Zhur. neorg. khim. 10 no.9:2203 S '65. (MIRA 18:10)

1. Institut obshchey i neorganicheskoy khimii imeni Kurnakova AN SSSR.

MOISEYEV, I.I.; BELOV, A.P., PER, G.Yu.

$\overline{\text{II}}$ -Hexenyl complex from propylene and palladium chloride.
Zhur. neorg. khim. 10 no.2:336-343 P '65. (MIRA 14:11)

1. Institut obshchey i neorganicheskoy khimii imeni Kurnakova
AN SSSR i Institut teknoy khimicheskoy tekhnologii imeni
Lomonosova. Submitted July 12, 1964.

MOISEYEV, I.I.; PESTRIKOV, S.V.

Palladium (0) complexes in butylene isomerization. Izv. AN
SSSR. Ser. khim. no.9:1717 '65. (MIRA 18:9)

1. Institut obshchey i neorganicheskoy khimii im. N.S.
Kurnakova AN SSSR.

MOISEYEV, I.I.; VARGAFIK, M.N.

Carbonium ions in the reactions of oxidation of olefins by
palladium chloride, Izv. AN SSSR, Ser. khim. no. 4: 753-760
'65. (MIRA 18:5)

1. Institut obshchey i neorganicheskoy khimii im. N.S. Kurnakova
AN SSSR.

FESTRIKOV, S.V.; MOISEYEV, I.I.

Equilibrium of the formation of π -complexes of palladium chloride
with 1-butylene in water. Izv. AN SSSR Ser. khim. no. 2: (9-16) 146.
(MIRA 18:2)

1. Institut obshchey i neorganicheskoy khimii im. N.S. Kurnakova
AN SSSR.

ZAKHAROVA, I.A.; MOISEYEV, I.I.

Cyclopropane derivative of palladium. Izv. AN SSSR. Ser. khim.
no.10:1914-1915 O '64. (MIRA 17:12)

1. Institut obshchey i neorganicheskoy khimii im. N.S. Kurnakova
AN SSSR.

MOISEYEV, I. I.; VARGAFTIK, M. N.; SYRKIN, Ya. K.

New π -allyl complex of palladium. Izv AN SSSR Ser Khim no. 4:
775 Ap '64. (MIRA 17:5)

π -Complex of palladium with triphenylcyclopropenyl. Ibid.:775--
776.

1. Institut tonkoy khimicheskoy tekhnologii im. M. V. Lomonosova
i Institut obshchey i neorganicheskoy khimii im. N. S.
Kurnakova AN SSSR.

MOISEYEV, I.I.; VARGAFNIK, M.N.; SYRKIN, Ya.K.

Kinetic stages of ethylene oxidation by palladium chloride
in aqueous solutions. Dokl. AN SSSR 153 no.1:140-143 N '63.
(MIRA 17:1)

1. Institut obshchey i neorganicheskoy khimii AN SSSR.
2. Chlen-korrespondent AN SSSR (for Syrkina).

MOISEYEV, I.I.; VARGAFIK, M.N.; SYRKIN, Ya.K.

Equilibrium of complex-forming process between palladium chloride
and ethylene in aqueous solutions. Dokl. AN SSSR 152 no.1:147-150
S '63. (MIRA 16:9)

1. Institut obshchey i neorganicheskoy khimii im. N.S. Kurnakova
AN SSSR. 2. Chlen-korrespondent AN SSSR (for Syrkin).
(Palladium compounds) (Ethylene)

NIKIFOROVA, A.V.; MOISEYEV, I.I.; SYRKIN, Ya.K.

Oxidation of alcohols with palladium salts in aqueous so-
lutions. Zhur.ob.khim. 33 no.10:3239-3242 0 '63.
(MIRA 16:11)

MOISEYEV, I.I.; BELOV, A.P.; SYRKIN, Ya.K.

Interaction between propylene and palladium chloride in acetic acid solutions. Izv.AN SSSR.Ser.khim. no.3:1527-1528 Ag '63.
(MIRA 1610)

1. Institut obshchey i neorganicheskoy khimii im. N.S.Kurnakova
AN SSSR i Institut tonkoy khimicheskoy tekhnologii im. Lomonosova.
(Propene) (Palladium chloride)

VARGAFTIK, M.N.; MOISEYEV, I.I.; SYRKIN, Ya.K.

Effect of chlorine ions on the rate of oxidation of ethylene by
palladium chloride in aqueous solutions. Izv. AN SSSR. Otd.khim.nauk
no.6:1147 Je '63. (MIRA 16:7)

1. Institut obshchey i neorganicheskoy khimii AN SSSR.
(Ethylene) (Oxidation) (Palladium chlorides)

MOISEYEV, I.I.; VARGAFIK, M.N.; SYRKIN, Ya.K.

Kinetic isotope effect of ethylene oxidation by palladium chloride.
Izv. AN SSSR. Otd.khim.nauk no.6:1144-1145 Je '63. (MIRA 16:7)

1. Institut tonkoy khimicheskoy tekhnologii imeni Lomonosova.
(Ethylene) (Oxidation) (Palladium compounds)

VARGAFIK, M.N.; MOISEYEV, I.I.; SYRKIN, Ya.K.

Kinetics of ethylene oxidation by palladium salts in
aqueous solutions. Dokl. AN SSSR 147 no.2:399-402
N '62. (MIRA 15:11)

1. Moskovskiy institut tonkoy khimicheskoy tekhnologii
im. M.V. Lomonosova. 2. Chlen-korrespondent AN SSSR
(for Syrkin).

(Ethylene) (Oxidation) (Palladium salts)

VARGAFNIK, M.N.; MOISEYEV, I.I.; SYRKIN, Ya.K.; YAKSHIN, V.V.

Formation of allyl esters in the reaction of higher olefins with
palladium chloride in solutions of anhydrous carboxylic acids.
Izv. AN SSSR. Otd.khim.nauk no.5:930-931 My '62. (MIRA 1:6)

1. Institut tonkoy khimicheskoy tekhnologii im. M.V.Lomonosova.
(Olefins) (Palladium chloride) (Esters)

VARGAFTIK, M.N.; MOISEYEV, I.I.; SYRKIN, Ya.K.

Kinetics of cyclohexane oxidation by palladium salts in aqueous solutions. Dokl. AN SSSR 139 No.6:1396-1399 Aug '61.

(MIRA 14:8)

1. Moskovskiy institut tonkoy khimicheskoy tekhnologii im. M.V. Lomonosova. 2. Chlen-korrespondent AN SSSR (for Syrkina).
(Cyclohexane) (Oxidation) (Palladium chloride)

KLIMENKO, N.M.; MOISAYEV, I.I.; SYRKIN, Ya.K.

Synthesis of a benzonitrile derivative of allylpalladium chloride.
Izv. AN SSSR. Otd.khim.nauk no.7:1355 J1 '61. (MIRA 14:7)

1. Institut tonkoy khimicheskoy tekhnologii im. M.V. Lomonosova.
(Palladium compounds)

81862

The Mechanism of the Reaction of Palladium Salts S/020/60/133/02/35/069
With Olefins in Hydroxyl-containing Solvents B016/B060

the other assumption of reaction (2) representing the first stage of decomposition, are held to be less probable. There are 4 references: 2 Soviet, 1 American, and 1 German.

ASSOCIATION: Moskovskiy institut tonkoy khimicheskoy tekhnologii im.
M. V. Lomonosova (Moscow Institute of Fine Chemical
Technology imeni M. V. Lomonosov)

SUBMITTED: April 23, 1960

Card 4/4

81862

The Mechanism of the Reaction of Palladium Salts S/020/60/133/02/35/068
With Olefins in Hydroxyl-containing Solvents B016/B060

oxidizing the palladium developed in the reaction, reaction (3) can be evidently used for the preparatory stage in the production of vinyl esters. The overall reaction in the presence of benzoquinone (see scheme) shows that also palladium is oxidized besides reaction (3). The authors' experiments further revealed that the above-mentioned complex reacts readily with alcohol and yields acetal as the main product (4). In the presence of p-benzoquinone, the reduction of PdCl_2 is accompanied by an oxidation of metallic Pd by way of ethylene in alcoholic solutions. This makes it possible for this reaction to be utilized in the direct production of acetals from olefins (see scheme). Also copper-salt solutions can be used as oxidizers in alcoholic solutions. The data obtained confirm the opinion that the decomposition of the π -complex in the hydroxyl-containing solvents takes place by way of the intermediate formation of vinyl compounds. The information supplied by the authors does not answer the question as to which of the two reactions (conversion of the π -complex into I or into II) represents the first stage of the decomposition. However, the assumption of conversion of II into a vinyl compound proceeding more quickly than the acidolysis of II, and

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The Mechanism of the Reaction of Palladium
Salts With Olefins in Hydroxyl-containing
Solvents

S/020/60/133/02/35/066
B016/B060

carbonyl compound attains 95-99%. It may be rather assumed that the charge of the nucleophilic particle attacking the C-atom of the olefin double bond does not play any essential part, and the addition of the HO^- ion takes place by reaction of the π -complex with the solvent molecules (2). Basic data on the mechanism of the decomposition of the π -complex can be obtained by studying the reaction between PdCl_2 and the olefins in nonaqueous solutions. The authors' experiments revealed that the $(\text{PdCl}_2 \cdot \text{C}_2\text{H}_4)_2$ complex synthesized by the method devised by M. S.

Kharash (Ref. 3), which reacts instantaneously even with atmospheric moisture, remains unaltered in a glacial acetic solution for even 10 days. This complex is rapidly decomposed in solutions of ethyl- as well as benzoyl alcohol and phenol. Experiments conducted by the authors further revealed that palladium chloride in acetic acid solutions containing sodium acetate is reduced by ethylene according to equation (3). The yield of vinyl acetate is 97% if referred to the reacted ethylene. The $(\text{PdCl}_2 \cdot \text{C}_2\text{H}_4)_2$ complex also reacts with sodium acetate in glacial acetic acid to form vinyl acetate. In the presence of substances capable of

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Moiseyev, I. I.

81862

S/020/60/133/02/35/068
B016/B060

5.3200

AUTHORS: Moiseyev, I. I., Vargaftik, M. N., Syrkin, Ya. K. Corresponding Member of the AS USSR

TITLE: The Mechanism of the Reaction of Palladium Salts With Olefins in Hydroxyl-containing Solvents

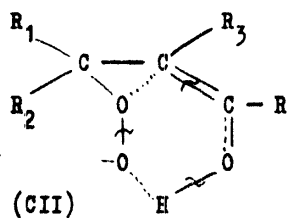
PERIODICAL: Doklady Akademii nauk SSSR, 1960, Vol. 133, No. 2, pp. 377-380

TEXT: In the authors' opinion, a participation of the HO^- ions in the reaction mentioned in the title appears to be little probable. Under the conditions mentioned in Ref. 1, the reaction of the π -complex with the HO^- ions is evidently accompanied by another reaction with the halide ions. This reaction leads to the formation of organohalogen compounds which are fairly stable under experimental conditions (the concentration of Br^- or Cl^- was 10^{10} - 10^{12} times higher than the OH^- concentration, Ref. 1). Nevertheless, such a scheme does not explain satisfactorily the high selectivity of the oxidation process in which the yield of the

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Mechanisms of Some Reactions With the Participation
of Peroxides

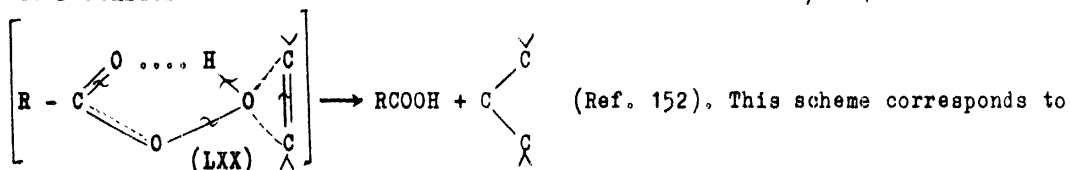
69633
S/074/60/029/04/01/005
B008/B014



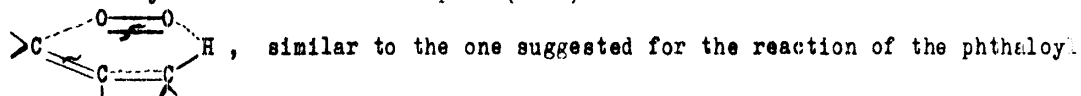
The following Soviet authors are mentioned: G. D. Lyubarskiy, M. Ya. Kagan, G. A. Razuvayev, V. M. Latyayeva, N. A. Prilezhayev, and O. A. Chaltykian. There are 1 table and 202 references, 29 of which are Soviet.

Mechanisms of Some Reactions With the Participation of Peroxides

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B008/B014



the data on the stereospecificity of the reaction according to Prilezhayev (Ref. 153). The relative yields of the oxidation products are obviously dependent on the distribution of electron density in the free radical. The extraordinarily smooth course of this reaction can, however, hardly be explained only by means of the electron effect in the free radical. The reaction probably takes place via a photo-activated cyclic transitional complex (XCII):



peroxide. A study of the formation of α -oxides in the alkaline oxidation of cis- and trans-benzalacetones, β -methyl trans-benzalacetophenone, and other derivatives with hydrogen peroxide showed that this is not a stereospecific reaction. The data concerning this reaction do not contradict the assumption that it proceeds via an activated complex of the type (CII):

Mechanisms of Some Reactions With the Participation of Peroxides

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B008/B014

confirmed by the high yield of corresponding acids. The resulting ester reacts according to either a molecular or a radical mechanism, depending on its structure. Besides molecular and radical reactions also the decomposition of the acyl peroxide itself is possible in the system acyl peroxide - alcohol. The ratio between the contributions of the molecular and the radical processes depends on temperature. The interaction between hydrogen peroxide and olefins may lead to different products, depending on the reaction conditions. (Refs. 10, 89, 90, 144-202). In the dark and without catalyst the reaction usually does not take place. The photosensitized reaction leads to the addition of hydrogen peroxide to the double bond. Besides this addition, decomposition of hydrogen peroxide in water and oxygen as well as oxidative side reactions are observable in this process. Oxidative side reactions proceed in the presence of carboxyl- and other functional groups in the olefin molecule. This reaction is likely to proceed like the hydrogenation of olefins that is catalyzed by acids. α -oxides are formed on interaction between olefins and per acids (reaction according to Prilezhayev - Ref. 147). All experimental data may be explained by means of the transition stage (LXX):

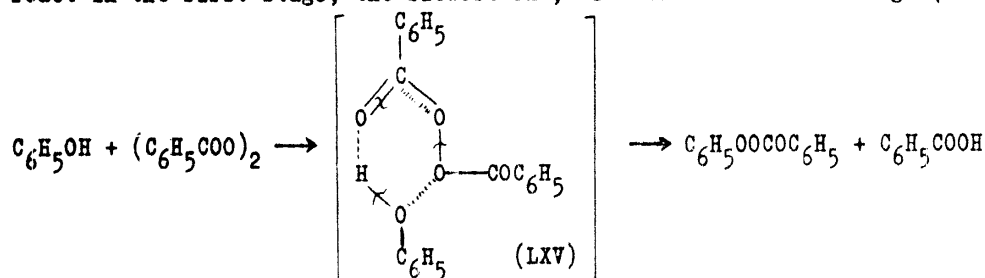
Card 5/7

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**Mechanisms of Some Reactions With the Participation
of Peroxides**

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B008/B014

solvent and cannot be separated. When studying the oxidation of phenols with acyl peroxides (Refs. 49, 126-143), the authors found that the reaction is not catalyzed by acids and bases. The reaction rate depends on the nature of the solvent in a very complicated manner. Besides polarity, also the existence of hydrogen bonds appears to be of some importance. For the rules governing this reaction the authors suggested a mechanism according to which benzoyl peroxide and phenol react in the first stage, the slowest one, via the intermediate stage (LXV):



It may be assumed that esters of per acids according to this scheme are also formed in reactions between acyl peroxides and alcohols. This assumption is

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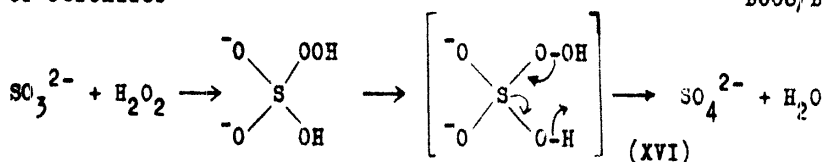
Mechanisms of Some Reactions With the Participation
of PeroxidesS/074/60/029/04/01/005
B008/B014

experimental data (p. 442). The corresponding hydroperoxides are formed during the reaction of hydrogen peroxide with alcohols (Refs. 41, 84-92). Also this reaction is accompanied by disruption of the alkyl-oxygen bond. The polarity of this bond is of great importance to the course of reaction. Presumably, the cyclic mechanism and mechanisms in which carbonium ions are formed compete with one another in this process. The thermal decomposition of per acids, esters of per acids, and acyl peroxides (Refs 49, 93-122) is accompanied by total or partial separation of CO_2 according to molecular mechanisms without free radicals participating in this reaction. It is often very difficult, sometimes even impossible, to distinguish the molecular reaction from the reaction of the radicals in the cage of the solvent. A few data obtained by means of optically active acyl-peroxides are indicative of a simultaneous occurrence of molecular and radical processes. In some cases, it is also possible that the individual processes compete with one another. The first stage of the decomposition reaction is presumably the decomposition of the molecule to form acyl-oxy radicals, as is the case with the noninitiated decomposition of benzoyl peroxide. The recombination of the radicals immediately after their formation is termed "primary recombination". In explaining the cage effect the authors refer to a paper by one of the authors mentioned in Ref. 125. This effect, which might be important to the mechanisms of several reactions in solutions, possibly consists in an interaction between two radicals or molecular fragments being in the cage, i.e., when they are surrounded by the

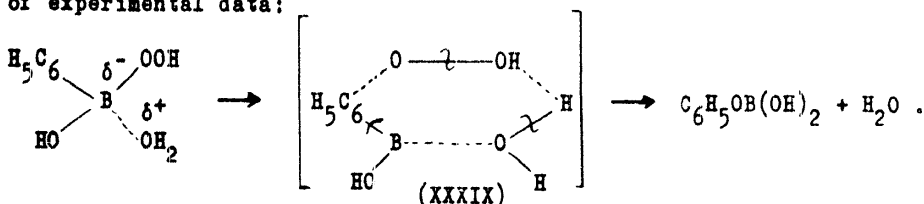
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Mechanisms of Some Reactions With the Participation of Peroxides

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B008/B014



The rearrangement of 9-decalylperbenzoate (Refs. 11, 41-81) indicated the intramolecular character of this reaction. The following scheme was set up on the basis of experimental data:



The elevation of the reaction rate is proportional to a pH of between 3 and 6. The corresponding acid and hydrogen peroxide are formed during the hydrolysis of a per acid (Refs 24, 70, 82). This reaction is accompanied by disruption of the acyl-oxygen bond. The reaction rate is proportional to the concentration of hydrogen ions, but not to the acidity of the solution. The mechanism previously suggested for several hydrolytic reactions of organic esters makes it possible to explain

6608 69633

5.3200

S/074/60/029/04/01/005
B008/B014AUTHORS: Syrkin, Ya. K., Moiseyev, I. I.TITLE: Mechanisms of Some Reactions With the Participation of Peroxides //PERIODICAL: Uspekhi khimii, 1960, Vol. 29, No. 4, pp. 425-469

TEXT: In the article under review, the authors describe their attempts to study a great number of reactions. They endeavored to find the mechanisms which correspond to the experiment or at least do not contradict it. They intended above all to clarify the role played by cyclic, activated complexes in reactions of peroxide compounds. First, they studied the oxidation of ketones and aldehydes with per acids and hydrogen peroxide (Refs. 1-40). It was found that the reaction is nearly always a second-order reaction, and is catalyzed by acids. The mechanisms under consideration had one feature in common, i.e., the formation of intermediates by addition of the hydrogen peroxide to the carbonyl bond. It seems to be reasonable to assume that several oxidation reactions of inorganic compounds also proceed via corresponding addition products. The authors established that the following scheme underlies the oxidation mechanism: X

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New Complexes of Palladium With Unsaturated Organic
Ligands

05893

SOV/78-4-11-46/50

of the bond with PdCl are made. A detailed report on further investigations in this field will follow. The authors thank V. I. Belova for the measurements of magnetic susceptibility, I. Yu. Kokoreva for the measurements of the dipole moments, and Yu. G. Borod'ko for recording the infrared spectra. There is 1 reference.

SUBMITTED: July 9, 1959

Card 2/2

5(2)

05893
SOV/78-4-11-46/50

AUTHORS:

Moiseyev, I. I., Fedorovskaya, E. A., Syrkin, Ya. K.

TITLE:

New Complexes of Palladium With Unsaturated Organic Ligands

PERIODICAL:

Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 11,
pp 2641-2642 (USSR)

ABSTRACT:

Palladium chloride reacts with allyl alcohols in acidified aqueous solution to form the compound C_3H_5PdCl . At temperatures between 15 and 20°, a yellow α -form develops which is stable up to 130°; at temperatures below 10°, the green β -form develops which already decomposes at 50°, dissolved in benzene even at 25°. The α -form is assumed to be identical with the compound $Pd_2Cl_2 \cdot C_6H_{10}$ as described in reference 1. Both forms are diamagnetic. The allyl group seems to be capable of forming a special type of complex compounds which is also proved by the reaction of palladium chloride with mesityl oxide. A diamagnetic yellow substance $PdCl(C_6H_5)_2$ develops which is stable up to 176°. Its infrared spectrum is discussed, and assumptions on the kind

Card 1/2

Mechanism of Organic Reactions Diene Synthesis'

SOV 21-27-11: 5

free mobility of 6 electrons in the cyclic transition complex with little expenditure of energy, however, not in the biradical. In the meeting of cyclic dienes and dienophiles spatial configurations are formed. Also in this case two possibilities exist. Formation of an endo- and an exo-configuration. Mainly endo-configuration is formed. Hitherto it has not been known why preference is given to endo-configuration. There exist several assumptions in this connection, however, each of them may only explain one fact, others are in contradiction to them. Woodward assumes the formation of an inter-molecular donor-acceptor binding (Ref 36) between the participants in diene synthesis (ionic linkage). Hudson and Robinson (Refs 36,37) assume the formation of a hybrid bond. Finally the authors arrived at the result that the double bonds play the most important part in the active complex. In the general case of diene synthesis it was therefore found that it takes place over a 6-cyclic active complex. Alder has expressed this fact in a similar way: for the reaction of the diene synthesis the presence of a hydrogen atom at a carbon atom with double bond is necessary and this double bond must be able to migrate. There are 3 tables and 54 references, 9 of which are Soviet.

Card 2/2

AUTHORS: Syrxin, Ya.K., Moiseyev, I.I. (Moscow, USSR) 307, 71-21-11-2, 3

TITLE: Mechanism of Organic Reactions (Mekhanizmy organicheskikh reaktsiy)
Diene Synthesis (Diyenovyy sintez)

PERIODICAL: Uspekhi khimii, 1958, Vol 27, Nr 11, pp 1321-1336 (USSR)

ABSTRACT: In this paper various hypotheses on the course of the diene synthesis are investigated. Two courses were assumed:
1) Formation of a transition complex with open side-chain (biradical), or biradical and subsequent cyclization.
2) Formation of a cyclic transition complex.
Computations of the activation energies necessary in any case have proved that the intermediate formation of a biradical and the subsequent cyclization are energetically impossible (attempt of dimerization of butadiene and formation of cyclohexane from butadiene and acetylene. The activation energy of the formation of cyclohexane is higher than that of the dimerization of butadiene) (Table 1). Experiments carried out by Eisler and Wassermann (Ref 3) confirmed the assumption of the cyclic transition complex. For the formation of a 6-cycle of two butadiene molecules the cis-form of diene is necessary which is possible only at the

Card 1/2

BELOV, A.P.; VARGAFIK, M.N.; MOISEYEV, I.I.

Bromination of π -allyl complexes of palladium. Izv. AN
SSSR. Ser. khim. no.8:1551-1552 Ag '64. (MIRA 1964)

1. Institut obshchey i neorganicheskoy khimii im. Kurnakova
AN SSSR i Institut tankoy khimicheskoy tekhnologii im.
Lomonosova.

The Mechanisms of Organic Reactions.
Hydrolysis and Esterification

100,74-27-6-4/6

these complexes play an important part in the reactions of hydrolysis. It must further be considered that in a number of cases it is possible that the reaction may develop by way of acyclic, open, and active complexes without the participation of carbon ions. There are 1 table and 61 references, 4 of which are Soviet.

1. Organic materials--Chemical reactions
2. Esters--Hydrolysis
3. Chemical reactions--Analysis

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The Mechanisms of Organic Reactions.
Hydrolysis and Esterification

SOV/1A-1-1-A 6

stages, (which are referred to here as II, A), etc., etc. is discussed. Also the reviews of L. S. Wintherwood (unpublished) and his collaborators (refs. 1, 2) are mentioned and discussed. Mention is also made of the articles by L. S. Wintherwood (Leydler) and L. S. Landkroner (Landkroner) (ref. 3): "Investigation of the Influence Exercised by the Solvent on the Hydrolysis of the Ethyl Acetate, Ethyl Propionate, Ethyl Butyrate, Propionamide, and Acrylamides of Various Acids in Aqueous and Acid Media". Next, also other pertinent transitional situations in hydrolysis reactions are discussed. Mention is made also of the method of nuclear-magnetic resonance (ref. 4) as well as of facts connected with the influence exercised by substituents in the benzene ring upon the kinetics of the hydrolysis of benzoate ester (refs. 5, 6), as also of the old terminology (ref. 7): attempt at explaining the rent observed on the alkyl-oxygen bond. Summary: The material discussed in this article shows that the hypotheses concerning unstable intermediate ions are as yet not sufficiently well founded and must in most cases be revised. Besides, the part played by cyclical active complexes in the reactions of hydrolysis has hitherto been underestimated. It is, however, certain that

Card 2/3

SOV/P-1-4-6

AUTHORS: Lytkin, A. I., Moiseyev, I. I. Moscow.

TITLE: The mechanism of esterification (mekhanizm organicheskikh reaktsiy) **Esterifikatsiya** (Gidroliz i esterifikatsiya)

PERIODICAL: Uspekhi khimii, 1958, 27, no. 6, pp. 1111-1112 (1111)

ABSTRACT: **Esterification** and hydrolysis are the parts of one and the same equilibrium. Therefore conclusions drawn with respect to the mechanism of one of these reactions refer also to the mechanism of the other. The character of the hydrolysis and of the kinetic law governing the reaction depend on the structure of the **ester** (which was subjected to hydrolysis) on the one hand, on the medium and on the catalyst on the other. This indicates the possibility of the existence of various different mechanisms. A. K. ... (...) and ... Ingold (Ingold ...) attempted ... to explain the experimental data for hydrolysis of **esters** by elementary processes. The various limiting cases are then enumerated. The energy of four of the total of eight of these

Card 1/3

[illegible]

20-3-33/59

The Mechanism Underlying the Hydration of Olefines in the Water Solutions of Strong Acids.

stage amounts to $-80-180+145+39+10 = -66$ kcal. For the solution of the question of the thermal effect of the second stage the knowledge of the formation heat of the C-ion from isobutylene and from proton is necessary. The C-ion proves to be more stable than the π -complex, and the transformation of the latter into the former takes place with an energy yield of more than 30 kcal. It is, however, necessary to assume that not the first, but the second stage is limited. Therefore these assumptions cannot be considered as proved. In the case of isobutylene the reaction process is possible because of the C-ion. This is, however, not possible since the lacking mutual transformation of 2-methylbutylene-2 and of 2-methylbutylene-1 in the case of insufficient dehydration eliminates the mechanisms in the presence of the C-ion as an intermediate substance which is in equilibrium with the initial substances. If the reaction passes through an activated complex-C-ion, the activation energies of isobutylene and propylene differ by about 15 kcal. In the experiment this difference amounts to 3,2 kcal. The authors consider a mechanism of this reaction in water solutions through a transition state (scheme I) to be more probable. They think it is essential that the proton-transfer-"stages" from H_3O^+ to the olefine molecule and the alcoxonium -ion-formation and H_3O^+ regeneration are combined in one single act. The consideration of the kinetic tasks is sometimes

Card 2/3

Moiseyev, I. I.

20-3-33/59

AUTHORS
TITLE

Moiseyev, I. I., Syrkin, Ya. K., Corresponding Member, AN USSR
The Mechanism Underlying the Hydration of Olefines in the Aqueous
Solutions of Strong Acids.
(O mekhanizme gidratatsii olefinov v vodnykh rastvorakh silnykh
kislot - Russian).

PERIODICAL

Doklady Akademii Nauk, 1957, Vol 115, Nr 3, pp 541-544 (U.S.S.R.)

ABSTRACT

The hydration of butylene, triptane, trimethylethylene, and methylcyclobutane in 1-5 M HNO_3 solutions is subject to an equation of first order. Between the observed velocity constant of the reaction k and the acidity function O^{H} of Hammett there is a dependence $\text{p-H}_0 = \lg k + \text{const}$, (1), the coefficient ρ being 1(0.98+1.11). The satisfying of this equation served as basis for the (here given) hydration mechanism of the olefines. The thermal effect of the first stage is composed of the following processes: a) proton desolvation -260, b) proton connection to the olefine under formation of a π -complex ω_x , c) "solvation" of the formed π -complex ω_x , d) dissolving of the water molecule + 10 cal bound to the proton. For the calculation of ω_x within the scale of the method MO IKA0 (meaning cannot be deciphered), the knowledge of the resonance integral of the π -binding between the atoms C and H in the π -complex is necessary. The π -value can be estimated by means of a measure analogous to that of Simonetta and Winstein. Approximative calculations of the energy yield in the proton connection to the double binding under the π -complex formation on the basis of various methods, give values of from 20 to 39 kcal. Thus the thermal effect of the first

Card 1/3

Subject : USSR/Chemistry

AID P - 3569

Card 1/1 Pub. 152 - 6/20

Authors : Flid, R. M. and I. I. Moiseyev

Title : Interaction of some components of catalysts during the liquid-phase hydration of acetylene

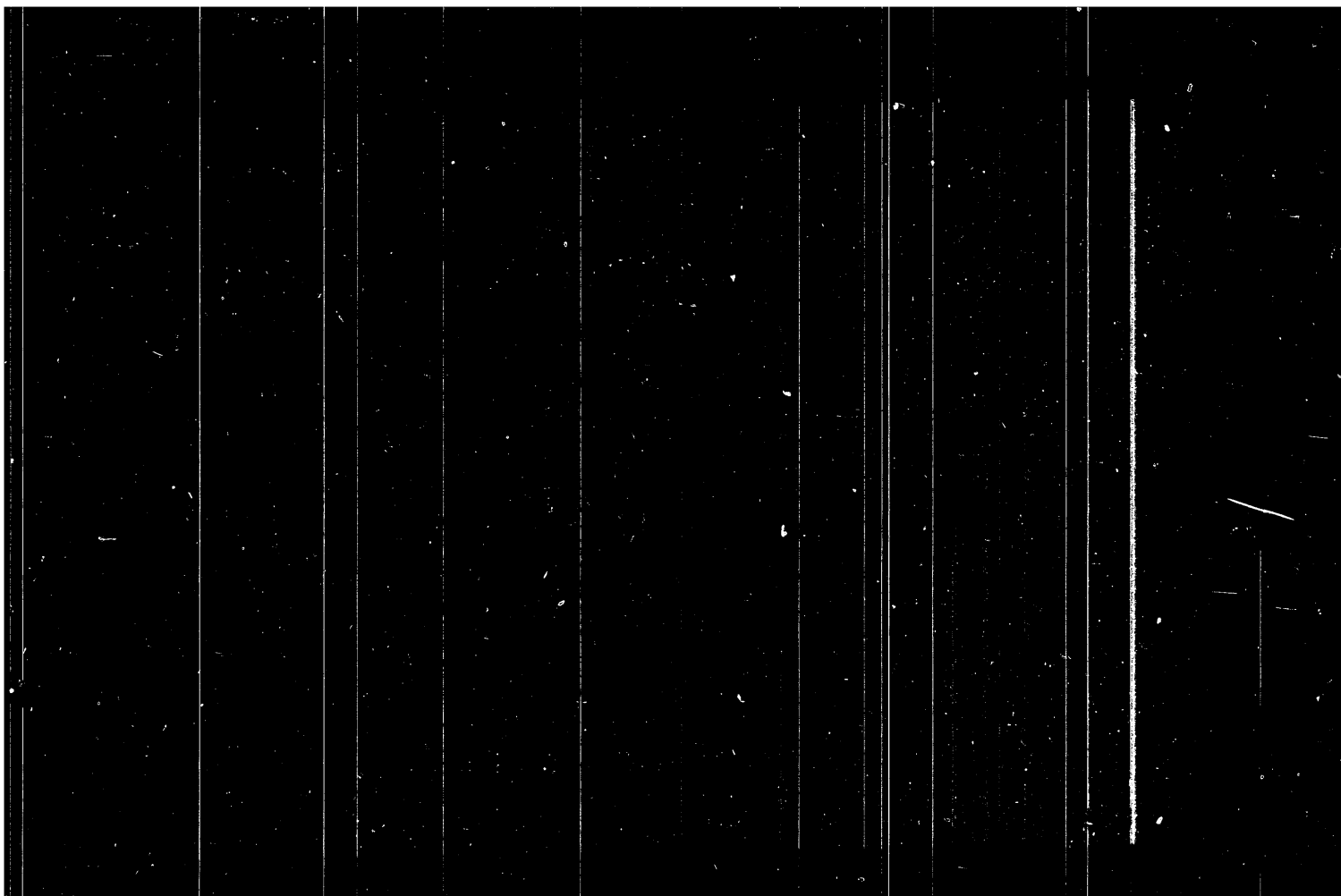
Periodical : Zhur. prikl. khim., 28, 711-717, 1955

Abstract : The reactions of mercury and of sulfates of mercury with ferrous and ferric sulfates have been studied. The experiments were carried out at 90-95°C in sulfuric acid. The effect of SO_4 -ion, aluminum sulfate and sodium sulfate on the voltage is determined, and the data compiled in tables. Six tables, 1 drawing, 10 references, 6 Russian (1946-1954).

Institution : Moscow Institute of Fine Chemical Technology im. Lomonosov. Chair of Technology of Basic Organic Synthesis

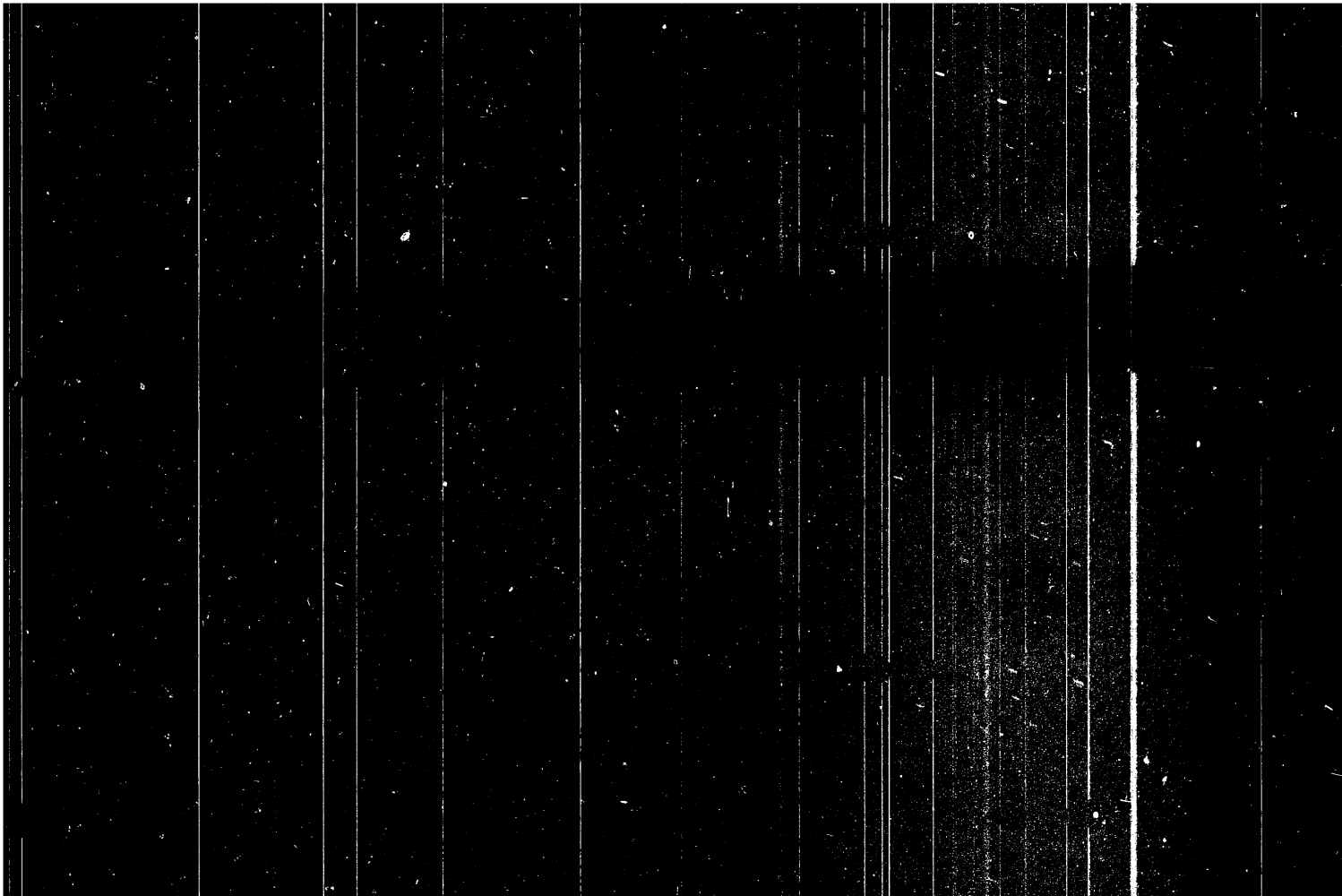
Submitted : Mr 30, 1953

APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6



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APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6

Abs Jour : Ref Zhur - Biol., No 10, 1958, No 43820

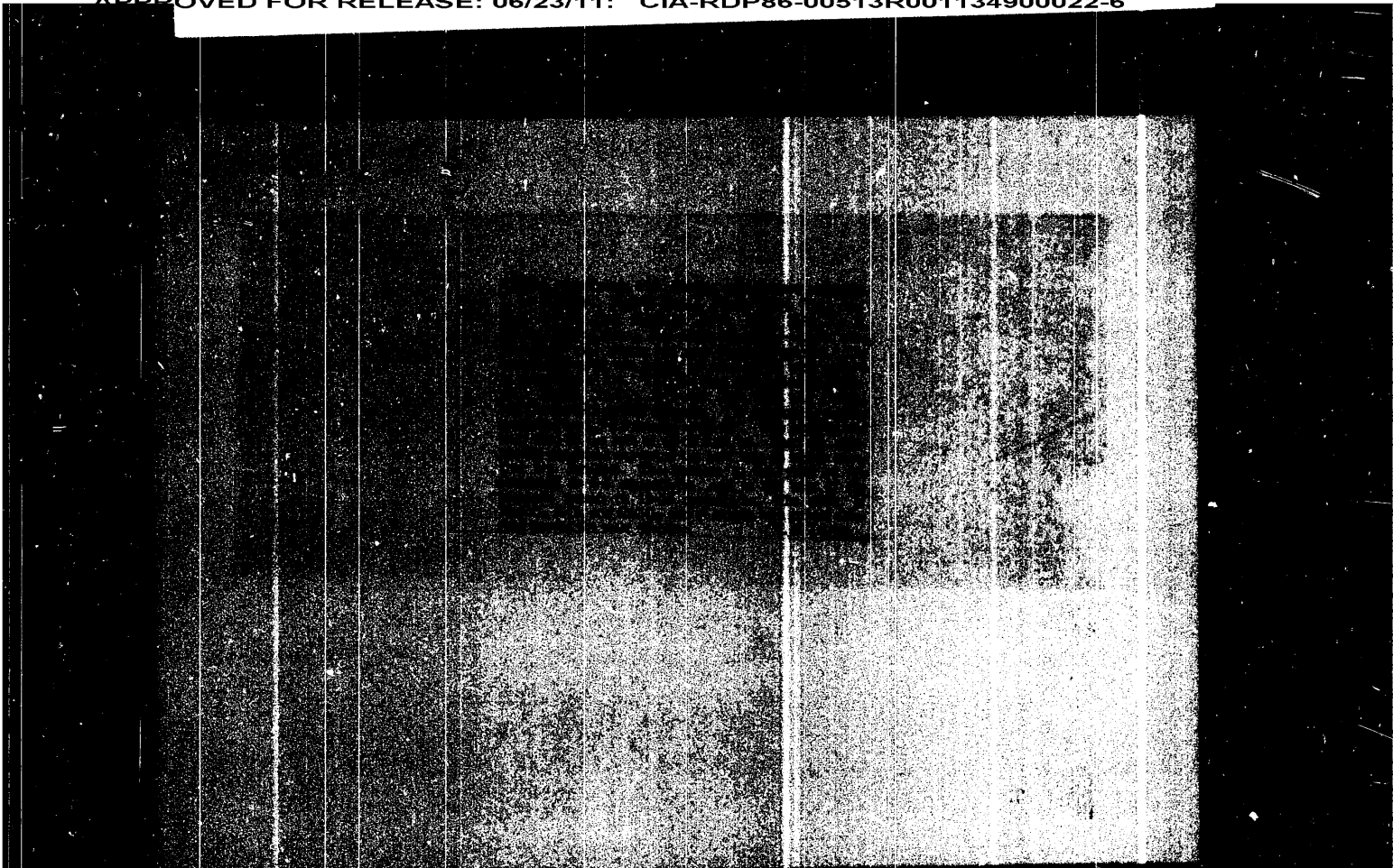
Author : Moiseyev I.G.
Inst : The Belorussian Institute of Agriculture
Title : Ways of Increasing the Fertility of Peat Podzolic Soils

Orig Pub : Zemledeliya, 1957, No 5, 59-63

Abstract ; The results of 8-year long field tests are expounded which were made by the Belorussian Institute of Agriculture on the strongly podzolized loess clay soil of Orshanskiy Rayon in Vitebskaya Oblast'. Peat-manure alkaline composts were tried out in double application during 9-field crop rotations with an overall dose of 50-85 t. per ha., and on individual plots with 28 t. of manure and 57 t. of the very same composts plus the addition of dolomite fertilizer and the mineral fertilizer $N_{35}P_{135}K_{165}$ [17]. The yield boosts obtained from 1 ha. over an average of 6 years were: winter rye 14.7 centners, perennial grass hay 56 centners and potatoes 141 centners. The absolute yield of winter rye obtained on the average from 1 ha. was 31.2 centners. --N.N. Sokolov

Card : 1/1

APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6



MOISEYEV, I. G.

Ratsional'noe primeneniye mestnykh udobrenii po dannym opytov i praktiki
peredovykh kolkhozov Belorusskoi SSR /Efficient application of local fertilizers
according to experimental data and practice of progressive collective farms in the
White Russian S.S.R./ Minsk, Gos. izd. BSSR, 1953. 46 p.

SO: Monthly List of Russian Accessions, Vol. 7 No. 1 April 1954.

APPROVED FOR RELEASE: 06/23/11: CIA-RDP86-00513R001134900022-6

MOISEYEV, I. G.

42433. K voprosu o melioratsii Podzolistykh pochv. izvestiya akad.
Nauk SSSR, No. 4, 1948, S. 87-97. Bibliogr: 8 Nazv.

L 10451-67

ACC NR: AP6022508

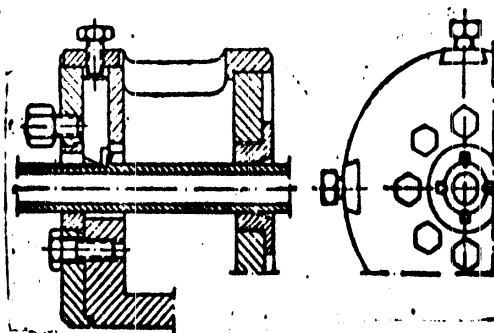


Fig. 1. Yoke for drawing longitudinal channels on the outer surface of pipes.

Orig. art. has: 3 graphs.

SUB CORR: 11/ SUBM DATE: none

Bimetal

Rev. 1/10/60

L 10451-67 EWT(m)/EWP(k)/EWP(t)/ETI IJP(c) JD/HW

ACC NR. AP6022508

SOURCE CODE: UR/0133/66/000/004/0348/0349

AUTHORS: Kaufman, M. Sh.; Shaykevich, S. A.; Kolmogorov, V. L.; Gleyberg, A. Z.;
 Alekhin, V. A.; Moiseyev, G. P.; Vostrikov, G. A.; Likhtenshteyn, D. Ye.; Gasilov,
 V. V.; Kuznetsov, S. N.; Borisov, L. M.

ORG: none

TITLE: Manufacture of two-layer pipes with continuous longitudinal channels between layers

SOURCE: Stal', no. 4, 1966, 348-349

TOPIC TAGS: pipe, steel, metal tube, metal forming

ABSTRACT: A method for manufacturing double layer steel Kh18N10T pipes with continuous longitudinal channels between the layers was developed. Two methods for the production of channels on the outer surface of the inner pipe were investigated--a rolling method and a cutting method. A schematic of the experimental installation is presented (see Fig. 1). It was found that both methods yielded pipes with smooth surfaces and uniform inner channels between the layers. The overall rate of pipe production, employing the cutting or drawing method, was 200 meters/hour. Double layer pipes having a diameter from 17 to 45 mm have been produced industrially. The following people took part in the experimental work: P. S. Ryshikov, N. A. Fedotovskiy, A. P. Nishkov, Ye. I. Tikhonov, and Ya. Z. Grinberg.

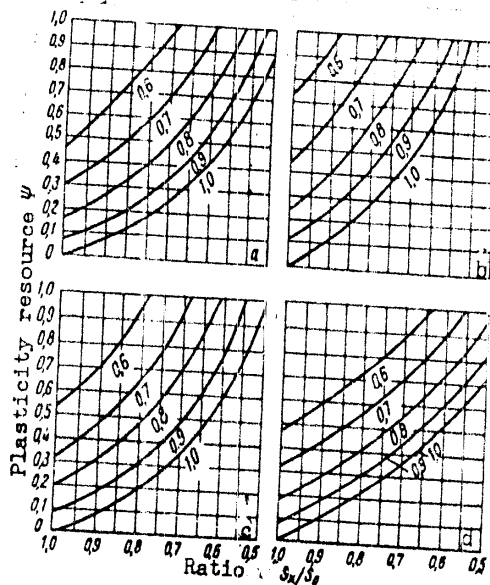
Card 1/2

UDC: 669.774.35

L 04814-67

ACC NR: AP6018388

Fig. 1. Use of the plasticity ψ during short-set drawing of pipes of steels 20 (a), 45 (b), 30KhGSA (c), and Kh18NiOT (d); numbers on the curves correspond to the change in pipe diameter d_1/d_0 as a result of drawing. S_x/S_0 = ratio of elongation coefficients.



production cycle, resulting in considerable savings in production costs. Orig. art. has: 1 table, 2 graphs, and 1 equation.

SUB CODE: 11/ SUBM DATE: none/ OTH REF: 001

Corr: 2/2 *gpl*